

Write your name here	
Surname	Other names
Pearson Edexcel Certificate Pearson Edexcel International GCSE	
Centre Number	Candidate Number
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<h1 style="margin: 0;">Chemistry</h1> <h2 style="margin: 5px 0 0 0;">Unit: KCH0/4CH0</h2> <h2 style="margin: 0 0 0 0;">Paper: 2C</h2>	
Wednesday 14 June 2017 – Morning Time: 1 hour	Paper Reference KCH0/2C 4CH0/2C
You must have: Calculator	Total Marks <div style="border: 1px solid black; width: 60px; height: 40px; margin: 0 auto;"></div>

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Show all the steps in any calculations and state the units.
- Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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THE PERIODIC TABLE

Period 1 2 3 4 5 6 7 0

Group

Period

4	He	Helium	2
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1	H	Hydrogen	1
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7	Li Lithium 3	9	Be Beryllium 4																	11	B Boron 5	12	C Carbon 6	14	N Nitrogen 7	16	O Oxygen 8	19	F Fluorine 9	20	Ne Neon 10				
23	Na Sodium 11	24	Mg Magnesium 12																	27	Al Aluminium 13	28	Si Silicon 14	31	P Phosphorus 15	32	S Sulfur 16	35.5	Cl Chlorine 17	40	Ar Argon 18				
39	K Potassium 19	40	Ca Calcium 20	45	Sc Scandium 21	48	Ti Titanium 22	51	V Vanadium 23	52	Cr Chromium 24	55	Mn Manganese 25	56	Fe Iron 26	59	Co Cobalt 27	59	Ni Nickel 28	63.5	Cu Copper 29	65	Zn Zinc 30	70	Ga Gallium 31	73	Ge Germanium 32	75	As Arsenic 33	79	Se Selenium 34	80	Br Bromine 35	84	Kr Krypton 36
86	Rb Rubidium 37	88	Sr Strontium 38	89	Y Yttrium 39	91	Zr Zirconium 40	93	Nb Niobium 41	96	Mo Molybdenum 42	99	Tc Technetium 43	101	Ru Ruthenium 44	103	Rh Rhodium 45	106	Pd Palladium 46	108	Ag Silver 47	112	Cd Cadmium 48	115	In Indium 49	119	Sn Tin 50	122	Sb Antimony 51	128	Te Tellurium 52	127	I Iodine 53	131	Xe Xenon 54
133	Cs Caesium 55	137	Ba Barium 56	139	La Lanthanum 57	179	Hf Hafnium 72	181	Ta Tantalum 73	184	W Tungsten 74	186	Re Rhenium 75	190	Os Osmium 76	192	Ir Iridium 77	195	Pt Platinum 78	197	Au Gold 79	201	Hg Mercury 80	204	Tl Thallium 81	207	Pb Lead 82	209	Bi Bismuth 83	210	Po Polonium 84	210	At Astatine 85	222	Rn Radon 86
223	Fr Francium 87	226	Ra Radium 88	227	Ac Actinium 89																	207	Pb Lead 82	209	Bi Bismuth 83	210	Po Polonium 84	210	At Astatine 85	222	Rn Radon 86				

Key

Relative atomic mass
Symbol
Name
Atomic number

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P 4 8 0 8 3 A 0 2 2 0

Answer ALL questions.

1 Sodium sulfate is a compound with many uses.

(a) The formula of the main compound used as the source of sodium sulfate is $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$

How many different elements are shown in this formula?

(1)

☐ A 2

☐ B 3

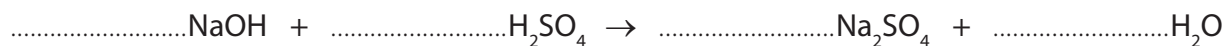
☐ C 4

☐ D 5

(b) Sodium sulfate can be made from sodium hydroxide and sulfuric acid.

Balance the equation for the reaction between sodium hydroxide and sulfuric acid.

(1)



(c) Sodium hydroxide is manufactured by the electrolysis of brine in the diaphragm cell.

Sulfuric acid is manufactured using the contact process.

The table contains some statements about these two processes.

Place ticks (✓) in the boxes to show the two correct statements.

(2)

brine is a solution of sodium chloride in water	
the temperature used in the contact process is greater than 1000 °C	
an equation for the contact process is $\text{SO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4$	
the reactions in the diaphragm cell are displacement reactions	
the catalyst used in the contact process is vanadium(V) oxide	

(Total for Question 1 = 4 marks)



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2 The diagram shows the positions of some elements in part of the Periodic Table.

						He	
Li		B		N		F	
	Mg		Si		S		Ar

(a) How many periods and groups are shown in this diagram?

(1)

	Periods	Groups
☐ A	2	4
☐ B	3	4
☐ C	2	8
☐ D	3	8

(b) How many elements shown in the diagram are noble gases?

(1)

- ☐ A 1
- ☐ B 2
- ☐ C 3
- ☐ D 4

(c) What is the formula of the compound formed between magnesium and fluorine?

(1)

- ☐ A MgF
- ☐ B Mg₂F
- ☐ C MgF₂
- ☐ D Mg₂F₂



(d) The table shows the percentage composition by mass of a sample of silicon.

Isotope	^{28}Si	^{29}Si	^{30}Si
Percentage (%)	92.2	4.70	3.10

Calculate the relative atomic mass of this sample of silicon.

Give your answer to one decimal place.

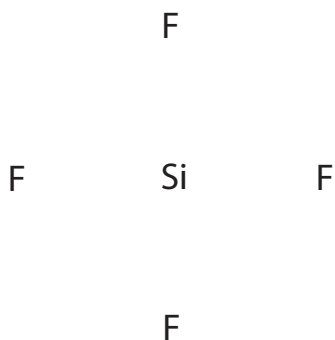
(2)

relative atomic mass =

(e) A molecule of silicon tetrafluoride (SiF_4) contains covalent bonds.

Draw a dot and cross diagram to show the outer electrons in this molecule.

(2)



- (f) The table shows the boiling points of some compounds containing silicon. All of these compounds contain covalent bonds.

Compound	Boiling point in °C
SiF ₄	–86
SiCl ₄	58
SiO ₂	2950

SiF₄ and SiCl₄ have simple molecular structures.

SiO₂ has a giant covalent structure.

- (i) Explain why the boiling point of SiCl₄ is greater than the boiling point of SiF₄ (2)

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- (ii) Explain why the boiling point of SiO₂ is very much greater than the boiling point of SiCl₄ (2)

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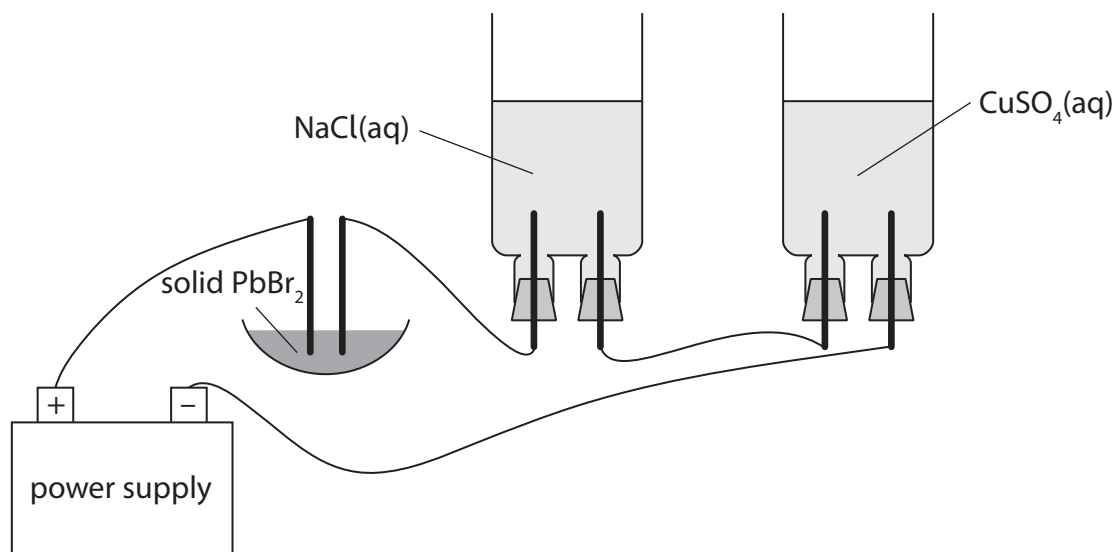
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(Total for Question 2 = 11 marks)



3 This apparatus can be used to investigate electrolysis.



(a) Name the particles that move through the connecting wires to form an electric current. (1)

(b) The electrodes are made of platinum, which is an inert metal.

State what is meant by the term **inert**.

(1)

(c) Explain why the electrolytic cell containing PbBr_2 needs to be heated before electrolysis can occur.

(2)

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- (d) When NaCl(aq) is electrolysed, two gases form at the positive electrode and one gas forms at the negative electrode.

The formulae of the species in NaCl(aq) are Na⁺, Cl⁻, H⁺, OH⁻ and H₂O.

- (i) Name the gases formed at each electrode.

(2)

positive electrode and

negative electrode

- (ii) Give ionic half-equations to show the formation of each gas.

(3)

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- (e) The ionic half-equation for one of the reactions in the cell containing copper(II) sulfate solution is



During the electrolysis, a charge of 0.040 faradays passes through this cell.

Calculate the mass of copper metal formed.

(2)

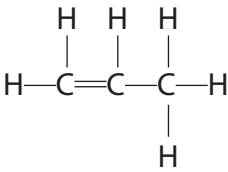
mass of copper = g

(Total for Question 3 = 11 marks)



- 4 Compounds containing C=C double bonds are used to manufacture alcohols such as ethanol and addition polymers such as PVC.

The table shows the formulae of some compounds containing C=C bonds.

A $\text{CH}_2=\text{CH}_2$	B 
C C_4H_8	D $\text{C}_2\text{H}_3\text{Cl}$

- (a) (i) Explain which three of these compounds are hydrocarbons.

(2)

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- (ii) Explain which compound is shown as a displayed formula.

(2)

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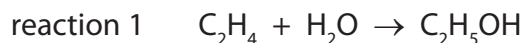
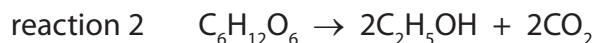
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- (b) Ethanol can be manufactured from compound A using reaction 1.
Ethanol can also be manufactured from glucose using reaction 2.

The equations for these reactions are



Give two advantages of using each reaction to manufacture ethanol.

(4)

reaction 1

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.....

.....

reaction 2

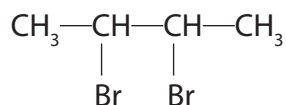
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- (c) Compound C reacts with bromine to form a product with this formula.



- (i) Use this formula to determine the name of compound C.

(2)

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- (ii) State the colour of the product formed when compound C reacts with bromine.

(1)

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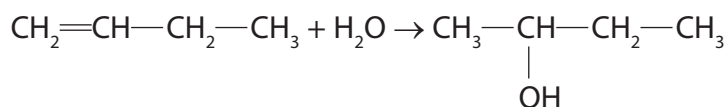


(d) Compound X is an isomer of compound C.

(i) Explain what is meant by the term **isomers**.

(2)

(ii) The equation for the conversion of compound X to an alcohol is



Place ticks (✓) in the boxes to show the two correct descriptions of this reaction.

(2)

addition	
dehydration	
hydration	
oxidation	
reduction	

(e) Compound D, chloroethene, can be used to manufacture the polymer PVC.

(i) State the full name of PVC.

(1)

(ii) Write an equation, containing displayed formulae, to represent the reaction that occurs in the manufacture of PVC.

(3)

(Total for Question 4 = 19 marks)



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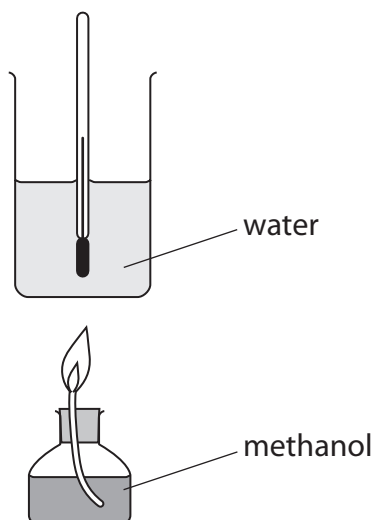
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- 5 A student uses this apparatus to find the increase in temperature of water when methanol, CH_3OH , is burned.



(a) There are several reasons why the increase in temperature is less than expected.

- (i) One reason is the incomplete combustion of methanol to form only carbon monoxide and water.

Write the chemical equation for this incomplete combustion.

(2)

(ii) State another reason why the increase in temperature is less than expected.

(1)

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(b) The student records these results.

mass of burner and methanol before combustion	84.7 g
mass of burner and methanol after combustion	83.2 g
mass of water	125 g
temperature of water at start	22 °C
temperature of water at end	58 °C

- (i) Calculate the heat energy change (Q), in joules, in this experiment using the expression

$$Q = m \times 4.2 \times \Delta T$$

where m is the mass of water in grams and ΔT represents the increase in temperature.
(2)

$Q =$ J

- (ii) The relative molecular mass of methanol is 32

Use this information and your value for Q to calculate the molar enthalpy change, ΔH , for the combustion of methanol.

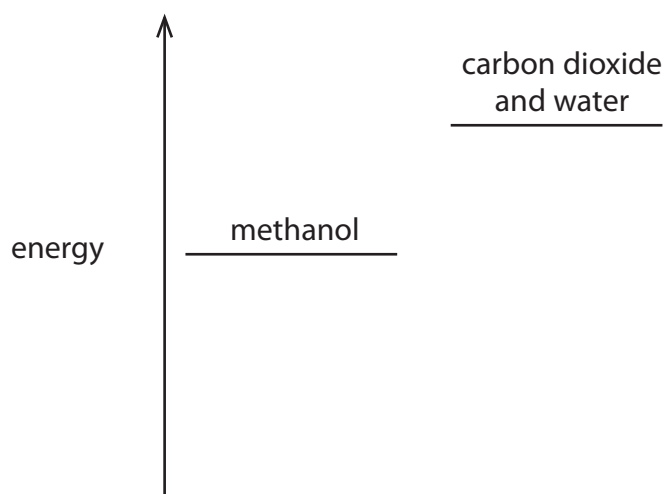
Give your answer in kJ/mol.

(4)

$\Delta H =$ kJ/mol



(iii) The student draws an energy level diagram for the complete combustion of methanol.



Identify the two mistakes in his diagram.

(2)

1

2

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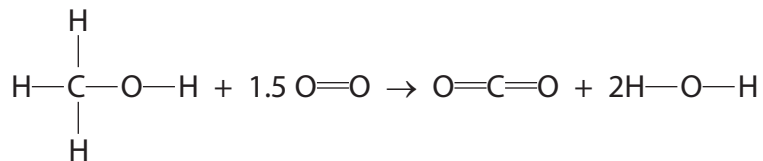
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(c) The student is given this table of average (mean) bond energies.

Bond	C—H	C—O	O—H	O=O	C=O
Average bond energy in kJ/mol	412	360	463	496	743

The equation for the complete combustion of methanol is



Use this equation and the information in the table to calculate another value for the molar enthalpy change, ΔH , for the combustion of methanol.

(4)

$\Delta H = \dots\dots\dots$ kJ/mol

(Total for Question 5 = 15 marks)

TOTAL FOR PAPER = 60 MARKS



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